

1-(2,4-Dichlorophenyl)-3-(4-methylphenyl)prop-2-en-1-one

Hoong-Kun Fun,^{a*} Samuel Robinson Jebas,^{a‡} P. S. Patil^b and S. M. Dharmaparakash^b

^aX-ray Crystallography Unit, School of Physics, Universiti Sains Malaysia, 11800

USM, Penang, Malaysia, and ^bDepartment of Studies in Physics, Mangalore

University, Mangalagangothri, Mangalore 574 199, India

Correspondence e-mail: hkfun@usm.my

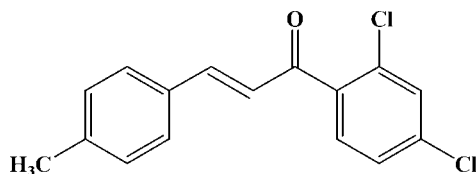
Received 19 April 2008; accepted 19 April 2008

Key indicators: single-crystal X-ray study; $T = 100$ K; mean $\sigma(\text{C}-\text{C}) = 0.002$ Å; R factor = 0.046; wR factor = 0.149; data-to-parameter ratio = 42.2.

The molecule of the title compound, $\text{C}_{16}\text{H}_{12}\text{Cl}_2\text{O}$, adopts an *E* configuration. The dihedral angle between the two benzene rings is $42.09(5)^\circ$. In the crystal structure, molecules are linked into a three-dimensional framework by weak $\text{C}-\text{H}\cdots\text{O}$ interactions and by $\text{C}-\text{H}\cdots\pi$ interactions involving the methylphenyl ring.

Related literature

For related literature, see: Agrinskaya *et al.* (1999); Gu *et al.* (2008); Patil *et al.* (2006); Patil, Dharmaparakash *et al.* (2007); Patil, Teh *et al.* (2007).



Experimental

Crystal data

$\text{C}_{16}\text{H}_{12}\text{Cl}_2\text{O}$

$M_r = 291.16$

Orthorhombic, *Pbca*

$a = 12.54850(1)$ Å

$b = 7.47750(1)$ Å

$c = 28.7764(3)$ Å

$V = 2700.13(3)$ Å³

$Z = 8$

Mo $K\alpha$ radiation

$\mu = 0.47$ mm⁻¹

$T = 100.0(1)$ K

$0.47 \times 0.39 \times 0.20$ mm

Data collection

Bruker SMART APEXII CCD

area-detector diffractometer

Absorption correction: multi-scan

(*SADABS*; Bruker, 2005)

$T_{\min} = 0.811$, $T_{\max} = 0.914$

50054 measured reflections

7296 independent reflections

4995 reflections with $I > 2\sigma(I)$

$R_{\text{int}} = 0.052$

Refinement

$R[F^2 > 2\sigma(F^2)] = 0.045$

$wR(F^2) = 0.148$

$S = 1.07$

7296 reflections

173 parameters

H-atom parameters constrained

$\Delta\rho_{\text{max}} = 0.64$ e Å⁻³

$\Delta\rho_{\text{min}} = -0.53$ e Å⁻³

Table 1

Hydrogen-bond geometry (Å, °).

$D-H\cdots A$	$D-H$	$H\cdots A$	$D\cdots A$	$D-H\cdots A$
$\text{C2}-\text{H2}\cdots\text{O1}^{\text{i}}$	0.93	2.55	3.4352 (15)	159
$\text{C8}-\text{H8}\cdots\text{O1}^{\text{ii}}$	0.93	2.55	3.1995 (14)	127
$\text{C11}-\text{H11}\cdots\text{Cg1}^{\text{iii}}$	0.93	2.81	3.5611 (13)	139
$\text{C14}-\text{H14}\cdots\text{Cg1}^{\text{iv}}$	0.93	2.93	3.7066 (13)	142

Symmetry codes: (i) $x + \frac{1}{2}, -y + \frac{3}{2}, -z$; (ii) $-x + \frac{1}{2}, y - \frac{1}{2}, z$; (iii) $-x + \frac{1}{2}, y + \frac{1}{2}, z$; (iv) $-x + 1, y - \frac{1}{2}, -z + \frac{1}{2}$. Cg1 is the centroid of the C10-C15 benzene ring.

Data collection: *APEX2* (Bruker, 2005); cell refinement: *APEX2*; data reduction: *SAINT* (Bruker, 2005); program(s) used to solve structure: *SHELXTL* (Sheldrick, 2008); program(s) used to refine structure: *SHELXTL*; molecular graphics: *SHELXTL*; software used to prepare material for publication: *SHELXTL* and *PLATON* (Spek, 2003).

FHK and SRJ thank the Malaysian Government and Universiti Sains Malaysia for the Science Fund grant No. 305/PFIZIK/613312. SRJ thanks the Universiti Sains Malaysia for the award of a post-doctoral research fellowship. PSP thanks the DRDO, Government of India, for a Senior Research Fellowship (SRF). This work is also supported by the Department of Science and Technology (DST), Government of India, under grant No. SR/S2/LOP-17/2006.

Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: CI2584).

References

- Agrinskaya, N. V., Lukoshkin, V. A., Kudryavtsev, V. V., Nosova, G. I., Solovskaya, N. A. & Yakimanski, A. V. (1999). *Phys. Solid State*, **41**, 1914–1917.
- Bruker (2005). *APEX2*, *SAINT* and *SADABS*. Bruker AXS Inc., Madison, Wisconsin, USA.
- Gu, B., Ji, W., Patil, P. S., Dharmaparakash, S. M. & Wang, H. T. (2008). *Appl. Phys. Lett.* **92**, 091118–091120.
- Patil, P. S., Dharmaparakash, S. M., Fun, H.-K. & Karthikeyan, M. S. (2006). *J. Cryst. Growth*, **297**, 111–116.
- Patil, P. S., Dharmaparakash, S. M., Ramakrishna, K., Fun, H.-K., Sai Santosh Kumar, R. & Rao, D. N. (2007). *J. Cryst. Growth*, **303**, 520–524.
- Patil, P. S., Teh, J. B.-J., Fun, H.-K., Razak, I. A. & Dharmaparakash, S. M. (2007). *Acta Cryst.* **E63**, o2122–o2123.
- Sheldrick, G. M. (2008). *Acta Cryst.* **A64**, 112–122.
- Spek, A. L. (2003). *J. Appl. Cryst.* **36**, 7–13.

‡ Permanent address: Department of Physics, Karunya University, Karunya Nagar, Coimbatore 641 114, India.

supplementary materials

Acta Cryst. (2008). E64, o936 [doi:10.1107/S1600536808011008]

1-(2,4-Dichlorophenyl)-3-(4-methylphenyl)prop-2-en-1-one

H.-K. Fun, S. R. Jebas, P. S. Patil and S. M. Dharmaprakash

Comment

Chalcone derivatives have been studied extensively owing to their fascinating, technologically relevant nonlinear optical properties (Gu *et al.*, 2008; Agrinskaya *et al.*, 1999; Patil *et al.*, 2006; Patil, Dharmaprakash *et al.*, 2007; Patil, Teh *et al.*, 2007).

The title molecule exhibits an E configuration with respect to the C8=C9 double bond [1.3424 (14) Å]; the C7—C8—C9—C10 torsion angle is 179.61 (11)°. The dihedral angle between the two benzene rings is 42.09 (5)°. The bond lengths and angles in the title molecule have normal values.

The crystal structure is stabilized by weak C—H···O intermolecular hydrogen bonding interactions (Table 1), which link the molecules into a three-dimensional framework (Fig. 2). In addition weak C—H··· π interactions involving the C10—C15 benzene ring (centroid Cg1) is observed.

Experimental

The title compound was synthesized by the condensation of p-tolualdehyde (0.01 mol) with 2,4-dichloroacetophenone (0.01 mol) in methanol (60 ml) in the presence of a catalytic amount of sodium hydroxide solution (5 ml, 30%). After stirring for 2 h, the contents of the flask were poured into ice-cold water (500 ml) and left to stand for 5 h. The resulting crude solid was collected by filtration and dried. Crystals suitable for single-crystal X-ray diffraction were grown by slow evaporation of an acetone solution at room temperature.

Refinement

H atoms were positioned geometrically [C—H = 0.93 (aromatic) or 0.96 Å (methyl)] and refined using a riding model, with $U_{\text{iso}}(\text{H}) = 1.5U_{\text{eq}}(\text{methyl C})$ or $1.2U_{\text{eq}}(\text{C})$.

Figures

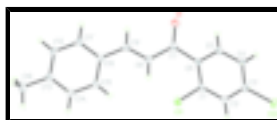


Fig. 1. The molecular structure of the title compound, showing 50% probability displacement ellipsoids and the atom numbering scheme.

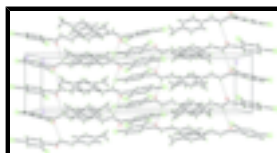


Fig. 2. The crystal packing of the title compound, viewed approximately along the *a* axis. Hydrogen bonds are shown as dashed lines.

1-(2,4-Dichlorophenyl)-3-(4-methylphenyl)prop-2-en-1-one

Crystal data

$C_{16}H_{12}Cl_2O$	$F_{000} = 1200$
$M_r = 291.16$	$D_x = 1.432 \text{ Mg m}^{-3}$
Orthorhombic, <i>Pbca</i>	Mo $K\alpha$ radiation
Hall symbol: -P 2ac 2ab	$\lambda = 0.71073 \text{ \AA}$
$a = 12.54850 (1) \text{ \AA}$	Cell parameters from 9639 reflections
$b = 7.47750 (1) \text{ \AA}$	$\theta = 2.8\text{--}34.6^\circ$
$c = 28.7764 (3) \text{ \AA}$	$\mu = 0.47 \text{ mm}^{-1}$
$V = 2700.13 (3) \text{ \AA}^3$	$T = 100.0 (1) \text{ K}$
$Z = 8$	Block, colourless
	$0.47 \times 0.39 \times 0.20 \text{ mm}$

Data collection

Bruker SMART APEXII CCD area-detector diffractometer	7296 independent reflections
Radiation source: fine-focus sealed tube	4995 reflections with $I > 2\sigma(I)$
Monochromator: graphite	$R_{\text{int}} = 0.052$
$T = 100.0(1) \text{ K}$	$\theta_{\text{max}} = 37.9^\circ$
φ and ω scans	$\theta_{\text{min}} = 1.4^\circ$
Absorption correction: multi-scan (SADABS; Bruker, 2005)	$h = -21 \rightarrow 21$
$T_{\text{min}} = 0.811$, $T_{\text{max}} = 0.914$	$k = -12 \rightarrow 12$
50054 measured reflections	$l = -48 \rightarrow 49$

Refinement

Refinement on F^2	Secondary atom site location: difference Fourier map
Least-squares matrix: full	Hydrogen site location: inferred from neighbouring sites
$R[F^2 > 2\sigma(F^2)] = 0.045$	H-atom parameters constrained
$wR(F^2) = 0.148$	$w = 1/[\sigma^2(F_o^2) + (0.078P)^2 + 0.0443P]$
$S = 1.07$	where $P = (F_o^2 + 2F_c^2)/3$
7296 reflections	$(\Delta/\sigma)_{\text{max}} = 0.001$
173 parameters	$\Delta\rho_{\text{max}} = 0.64 \text{ e \AA}^{-3}$
Primary atom site location: structure-invariant direct methods	$\Delta\rho_{\text{min}} = -0.53 \text{ e \AA}^{-3}$
	Extinction correction: none

Special details

Experimental. The data was collected with the Oxford Cyrosystem Cobra low-temperature attachment.

Geometry. All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

Refinement. Refinement of F^2 against ALL reflections. The weighted R -factor wR and goodness of fit S are based on F^2 , conventional R -factors R are based on F , with F set to zero for negative F^2 . The threshold expression of $F^2 > \sigma(F^2)$ is used only for calculating R -factors(gt) etc. and is not relevant to the choice of reflections for refinement. R -factors based on F^2 are statistically about twice as large as those based on F , and R -factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (\AA^2)

	<i>x</i>	<i>y</i>	<i>z</i>	$U_{\text{iso}}^*/U_{\text{eq}}$
C11	0.46470 (2)	0.76088 (4)	0.057148 (10)	0.02125 (8)
C12	0.36980 (3)	0.56187 (5)	-0.116828 (9)	0.02538 (9)
O1	0.13288 (7)	0.72551 (12)	0.08499 (3)	0.02032 (17)
C1	0.37202 (9)	0.67308 (15)	0.01850 (4)	0.0158 (2)
C2	0.40554 (10)	0.64861 (15)	-0.02709 (4)	0.0177 (2)
H2	0.4759	0.6696	-0.0356	0.021*
C3	0.33080 (10)	0.59184 (15)	-0.05947 (4)	0.0177 (2)
C4	0.22518 (9)	0.56292 (16)	-0.04764 (4)	0.0183 (2)
H4	0.1760	0.5265	-0.0699	0.022*
C5	0.19470 (9)	0.58948 (15)	-0.00210 (4)	0.0173 (2)
H5	0.1238	0.5715	0.0060	0.021*
C6	0.26690 (9)	0.64245 (14)	0.03220 (3)	0.01531 (19)
C7	0.22310 (9)	0.66606 (15)	0.08060 (4)	0.0165 (2)
C8	0.28861 (9)	0.60913 (15)	0.12039 (4)	0.0168 (2)
H8	0.3512	0.5458	0.1151	0.020*
C9	0.25956 (10)	0.64712 (15)	0.16419 (4)	0.0168 (2)
H9	0.1962	0.7100	0.1680	0.020*
C10	0.31720 (9)	0.59981 (15)	0.20655 (4)	0.01556 (19)
C11	0.27256 (10)	0.64353 (15)	0.24970 (4)	0.0179 (2)
H11	0.2075	0.7031	0.2508	0.021*
C12	0.32405 (10)	0.59919 (16)	0.29076 (4)	0.0189 (2)
H12	0.2929	0.6294	0.3190	0.023*
C13	0.42150 (10)	0.51030 (16)	0.29039 (4)	0.0182 (2)
C14	0.46702 (9)	0.46835 (16)	0.24731 (4)	0.0185 (2)
H14	0.5325	0.4102	0.2463	0.022*
C15	0.41598 (9)	0.51215 (16)	0.20617 (4)	0.0177 (2)
H15	0.4476	0.4831	0.1779	0.021*
C16	0.47564 (11)	0.45835 (18)	0.33516 (4)	0.0251 (3)
H16A	0.4802	0.5608	0.3552	0.038*
H16B	0.5460	0.4149	0.3286	0.038*
H16C	0.4352	0.3662	0.3503	0.038*

Atomic displacement parameters (\AA^2)

U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
----------	----------	----------	----------	----------	----------

supplementary materials

C11	0.01778 (15)	0.02973 (16)	0.01623 (13)	-0.00376 (10)	-0.00292 (9)	-0.00163 (10)
C12	0.02586 (17)	0.03750 (18)	0.01277 (12)	0.00235 (12)	0.00091 (10)	-0.00154 (10)
O1	0.0168 (4)	0.0256 (4)	0.0185 (4)	0.0022 (3)	-0.0001 (3)	-0.0017 (3)
C1	0.0158 (5)	0.0171 (5)	0.0145 (4)	0.0009 (4)	-0.0025 (4)	-0.0003 (3)
C2	0.0158 (5)	0.0216 (5)	0.0157 (4)	0.0016 (4)	0.0004 (4)	0.0001 (4)
C3	0.0213 (6)	0.0200 (5)	0.0117 (4)	0.0020 (4)	-0.0010 (4)	0.0007 (3)
C4	0.0183 (5)	0.0210 (5)	0.0156 (4)	-0.0001 (4)	-0.0043 (4)	-0.0007 (4)
C5	0.0166 (5)	0.0198 (5)	0.0154 (4)	-0.0007 (4)	-0.0014 (4)	0.0005 (4)
C6	0.0177 (5)	0.0155 (4)	0.0127 (4)	0.0013 (4)	-0.0010 (4)	0.0012 (3)
C7	0.0191 (5)	0.0152 (4)	0.0153 (4)	-0.0006 (4)	0.0005 (4)	-0.0008 (3)
C8	0.0188 (5)	0.0172 (5)	0.0144 (4)	0.0018 (4)	0.0001 (4)	0.0001 (3)
C9	0.0170 (5)	0.0173 (5)	0.0159 (4)	-0.0006 (4)	0.0003 (4)	-0.0001 (3)
C10	0.0177 (5)	0.0163 (4)	0.0127 (4)	-0.0004 (4)	0.0012 (4)	-0.0001 (3)
C11	0.0180 (5)	0.0200 (5)	0.0156 (4)	0.0021 (4)	0.0006 (4)	-0.0008 (4)
C12	0.0222 (6)	0.0218 (5)	0.0126 (4)	0.0004 (4)	0.0013 (4)	-0.0008 (4)
C13	0.0204 (6)	0.0184 (5)	0.0158 (5)	-0.0024 (4)	-0.0019 (4)	0.0007 (4)
C14	0.0168 (5)	0.0207 (5)	0.0179 (5)	0.0013 (4)	0.0002 (4)	0.0001 (4)
C15	0.0175 (5)	0.0209 (5)	0.0149 (4)	0.0013 (4)	0.0009 (4)	-0.0002 (4)
C16	0.0284 (7)	0.0290 (6)	0.0179 (5)	0.0016 (5)	-0.0058 (5)	0.0011 (4)

Geometric parameters (Å, °)

C11—C1	1.7380 (11)	C9—C10	1.4609 (15)
C12—C3	1.7360 (11)	C9—H9	0.93
O1—C7	1.2228 (14)	C10—C11	1.4008 (15)
C1—C2	1.3895 (15)	C10—C15	1.4022 (17)
C1—C6	1.3957 (16)	C11—C12	1.3871 (16)
C2—C3	1.3887 (16)	C11—H11	0.93
C2—H2	0.93	C12—C13	1.3919 (18)
C3—C4	1.3854 (17)	C12—H12	0.93
C4—C5	1.3795 (15)	C13—C14	1.4006 (16)
C4—H4	0.93	C13—C16	1.5072 (16)
C5—C6	1.3971 (15)	C14—C15	1.3855 (16)
C5—H5	0.93	C14—H14	0.93
C6—C7	1.5077 (15)	C15—H15	0.93
C7—C8	1.4725 (15)	C16—H16A	0.96
C8—C9	1.3424 (14)	C16—H16B	0.96
C8—H8	0.93	C16—H16C	0.96
C2—C1—C6	122.08 (10)	C10—C9—H9	116.6
C2—C1—C11	116.83 (9)	C11—C10—C15	118.00 (10)
C6—C1—C11	120.90 (8)	C11—C10—C9	119.02 (10)
C3—C2—C1	118.00 (11)	C15—C10—C9	122.98 (10)
C3—C2—H2	121.0	C12—C11—C10	120.88 (11)
C1—C2—H2	121.0	C12—C11—H11	119.6
C4—C3—C2	121.92 (10)	C10—C11—H11	119.6
C4—C3—C12	118.90 (9)	C11—C12—C13	121.12 (10)
C2—C3—C12	119.15 (10)	C11—C12—H12	119.4
C5—C4—C3	118.45 (10)	C13—C12—H12	119.4
C5—C4—H4	120.8	C12—C13—C14	118.17 (10)

C3—C4—H4	120.8	C12—C13—C16	120.83 (10)
C4—C5—C6	122.15 (11)	C14—C13—C16	120.99 (11)
C4—C5—H5	118.9	C15—C14—C13	121.00 (11)
C6—C5—H5	118.9	C15—C14—H14	119.5
C1—C6—C5	117.38 (10)	C13—C14—H14	119.5
C1—C6—C7	125.90 (10)	C14—C15—C10	120.83 (10)
C5—C6—C7	116.70 (10)	C14—C15—H15	119.6
O1—C7—C8	122.79 (10)	C10—C15—H15	119.6
O1—C7—C6	118.40 (10)	C13—C16—H16A	109.5
C8—C7—C6	118.75 (10)	C13—C16—H16B	109.5
C9—C8—C7	121.15 (11)	H16A—C16—H16B	109.5
C9—C8—H8	119.4	C13—C16—H16C	109.5
C7—C8—H8	119.4	H16A—C16—H16C	109.5
C8—C9—C10	126.71 (11)	H16B—C16—H16C	109.5
C8—C9—H9	116.6		
C6—C1—C2—C3	0.00 (17)	C5—C6—C7—C8	141.55 (11)
C11—C1—C2—C3	-175.09 (9)	O1—C7—C8—C9	-11.25 (18)
C1—C2—C3—C4	1.20 (17)	C6—C7—C8—C9	171.61 (10)
C1—C2—C3—C12	179.44 (9)	C7—C8—C9—C10	-179.61 (11)
C2—C3—C4—C5	-0.88 (17)	C8—C9—C10—C11	-177.16 (12)
C12—C3—C4—C5	-179.12 (9)	C8—C9—C10—C15	2.91 (19)
C3—C4—C5—C6	-0.67 (17)	C15—C10—C11—C12	-0.83 (17)
C2—C1—C6—C5	-1.44 (17)	C9—C10—C11—C12	179.24 (11)
C11—C1—C6—C5	173.45 (9)	C10—C11—C12—C13	0.09 (19)
C2—C1—C6—C7	-179.70 (10)	C11—C12—C13—C14	0.69 (18)
C11—C1—C6—C7	-4.80 (16)	C11—C12—C13—C16	-178.42 (11)
C4—C5—C6—C1	1.79 (17)	C12—C13—C14—C15	-0.73 (18)
C4—C5—C6—C7	-179.80 (10)	C16—C13—C14—C15	178.38 (11)
C1—C6—C7—O1	142.55 (12)	C13—C14—C15—C10	-0.01 (18)
C5—C6—C7—O1	-35.72 (15)	C11—C10—C15—C14	0.79 (17)
C1—C6—C7—C8	-40.19 (15)	C9—C10—C15—C14	-179.29 (11)

Hydrogen-bond geometry (Å, °)

<i>D</i> —H \cdots <i>A</i>	<i>D</i> —H	H \cdots <i>A</i>	<i>D</i> \cdots <i>A</i>	<i>D</i> —H \cdots <i>A</i>
C2—H2 \cdots O1 ⁱ	0.93	2.55	3.4352 (15)	159
C8—H8 \cdots O1 ⁱⁱ	0.93	2.55	3.1995 (14)	127
C11—H11 \cdots Cg1 ⁱⁱⁱ	0.93	2.81	3.5611 (13)	139
C14—H14 \cdots Cg1 ^{iv}	0.93	2.93	3.7066 (13)	142

Symmetry codes: (i) $x+1/2, -y+3/2, -z$; (ii) $-x+1/2, y-1/2, z$; (iii) $-x+1/2, y+1/2, z$; (iv) $-x+1, y-1/2, -z+1/2$.

Fig. 1

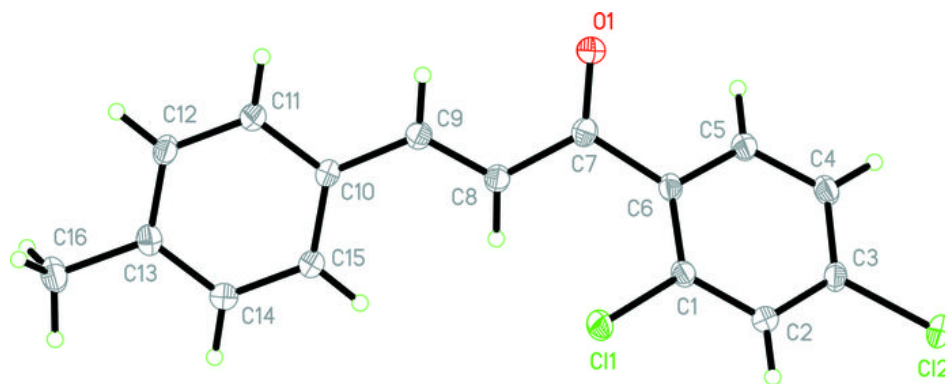


Fig. 2

